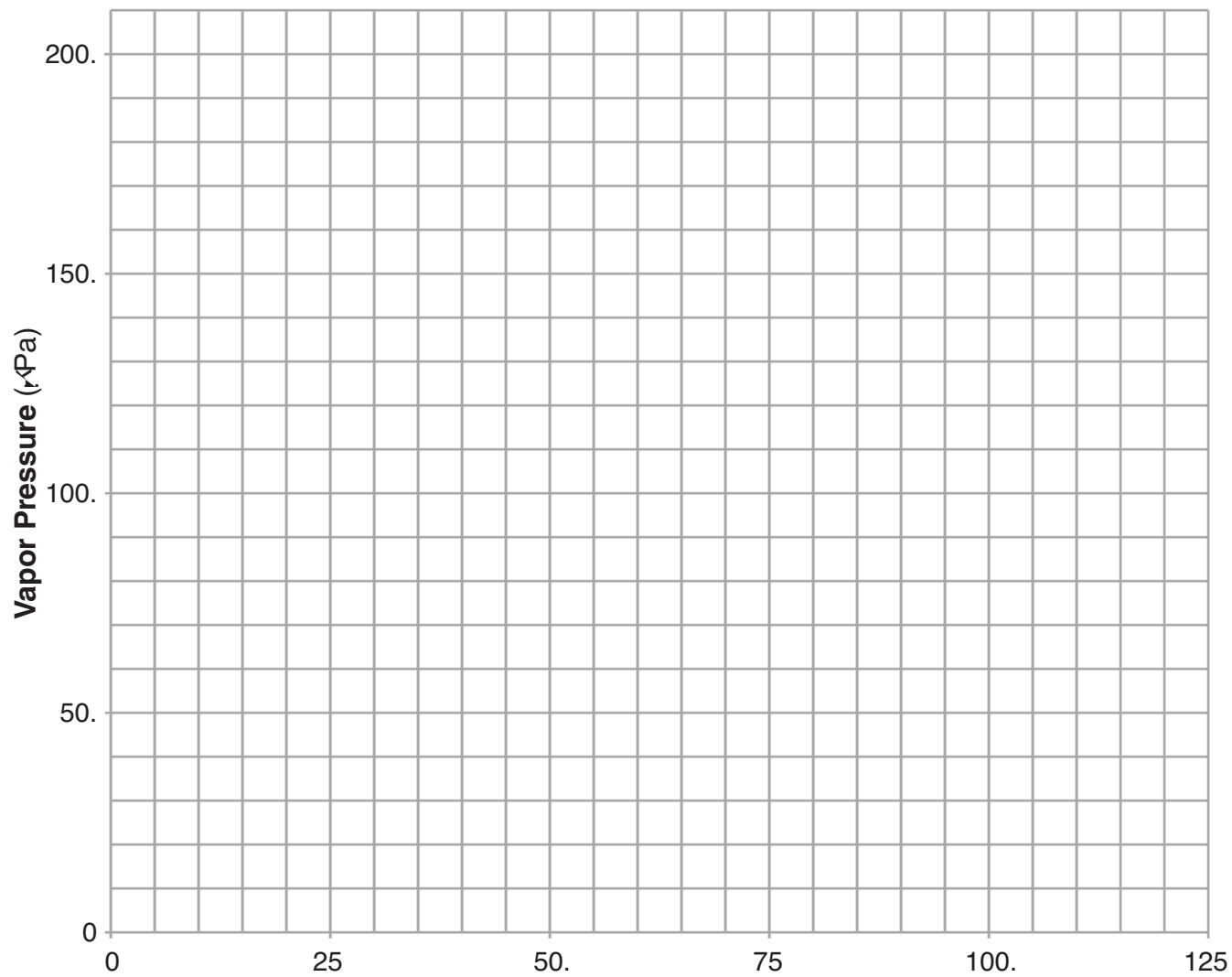




Ions That Form <i>Soluble</i> Compounds	Exceptions
$\text{Li}^+$ $\text{Na}^+$ $\text{K}^+$	
$\text{NH}_4^+$	
$\text{NO}_3^-$ $\text{ClO}_4^-$	



**Table H**  
**Vapor Pressure of Four Liquids**



**Table I**  
**Heats of Reaction at 101.3 kPa and 298 K**

Reaction	$\Delta H$ (kJ)*
<div style="text-align: center; margin-top: 20px;"> </div>	

**Table J**  
**Activity Series\*\***


**Table K**  
**Common Acids**

**Table N**

**Table L**  
**Common Bases**

## Table O

Name	General Formula	Examples	
		Name	Structural Formula
/ / / /		/ / / /	
/ / / /		/ / / /	
/ / / /		/ / / /	

**Table R**  
**Organic Functional Groups**

<b>Class of Compound</b>	<b>Functional Group</b>	<b>General Formula</b>	<b>Example</b>
		$R$	
		$R$	
		$R$	
		$R$	
		$R$	
		$R$	
		$R$	
		$R$	





**Table S**  
**Properties of Selected Elements**

Atomic Number	Symbol	Name	First Ionization Energy	Electro-negativity	Melting Point	Boiling Point	Density**	Atomic Radius
1								
2								
3								
4								
5								
6								
7								
8								
9								
10								
11								
12								
13								
14								
15								
16								
17								



**Table T**  
**Important Formulas and Equations**

<b>Density</b>	$\rho = \frac{m}{V}$
<b>Mole Calculations</b>	$n = \frac{m}{M}$
<b>Percent Error</b>	$\% \text{ Error} = \frac{ \text{Experimental} - \text{Theoretical} }{\text{Theoretical}} \times 100$
<b>Percent Composition</b>	$\% \text{ Composition} = \frac{\text{mass of element}}{\text{molar mass of compound}} \times 100$
<b>Concentration</b>	$M = \frac{n}{V}$
	$m = \rho \times V$
<b>Combined Gas Law</b>	$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
<b>Titration</b>	$M_A V_A = M_B V_B$
<b>Heat</b>	$Q = C \Delta T$ $H = C \times m \times \Delta T$ $H = H_f + H_v + H_c$
<b>Temperature</b>	$T(^{\circ}\text{C}) = T(^{\circ}\text{F}) - 32$